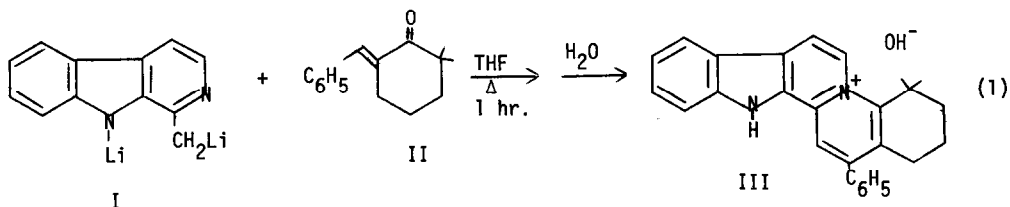


NEW BASE-CATALYZED PREPARATION OF CERTAIN QUINOLIZINIUM SALTS

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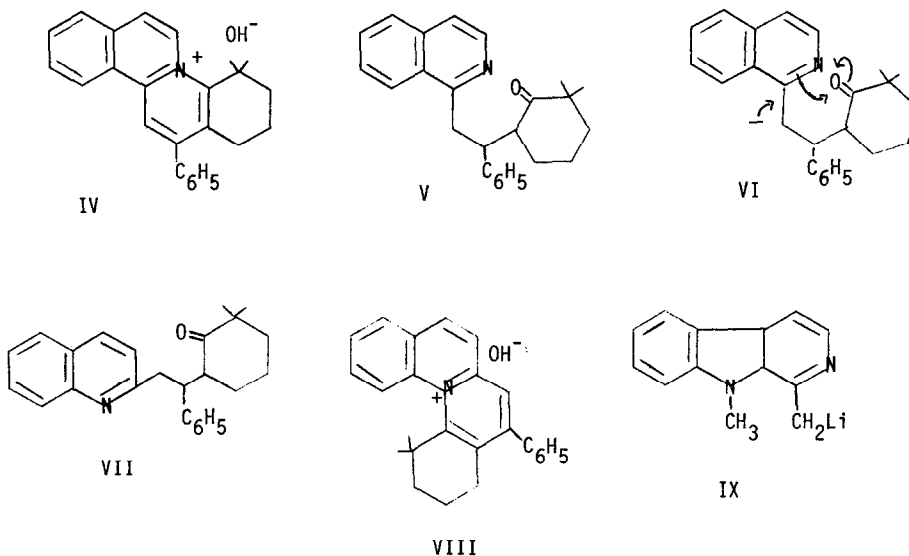
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As part of a continuing study of the chemistry of lithiomethylated nitrogenous heterocycles,¹ a new remarkably facile synthesis of certain quinolizinium salts has been found. The one-pot reactions consisting of a conjugate addition of such carbanions with α,β -unsaturated ketones followed by spontaneous in situ oxidative cyclizations are illustrated by the reaction of dilithioharman (I) with cyclohexanone II to afford indoloquinolizinium salt III² (49%)(eq. 1). To our knowledge, such cyclizations are the first ones effected by basic rather than acidic reagents.³



Similarly, 1-lithiomethylisoquinoline has been reacted with II in the presence of excess lithium diisopropylamide to afford benzoquinolizinium salt IV² (40-45%). Interestingly, this latter reaction in the absence of excess base fails to give cyclized product; instead, only the open-chain product V² is obtained (86%). These results suggest that the cyclization might proceed via a process illustrated by VI where a new α -methyl carbanion is formed which indirectly attacks the carbonyl (or the corresponding enolate anion) as shown. Loss of a "hydride" and dehydration of the alcohol or lithioalkoxide would complete the process.

In contrast, reaction of 2-lithiomethylquinoline with II gives the open-chain product VII² which, to date, has resisted cyclization to VIII. Perhaps such a cyclization is precluded in this case by peri-hydrogen interaction since it is nearly impossible to make a space-filling model of VIII.



At first glance, the above conjugate additions appear to be similar to the previously reported reaction of lithio-N-methylharman (IX) with isopropoxymethylene cyclohexanone to give sempervirine methochloride.⁴ However, this sequence involved a 1,2-addition to the ketone followed by an acid-catalyzed cyclization.

An active study is in progress to assess the types of lithiomethylated heterocycles and α,β -unsaturated carbonyl compounds that will undergo such condensations and oxidative cyclizations.

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References

1. Previous paper: E.M. Kaiser and W.R. Thomas, *J. Org. Chem.*, **39**, 2659 (1974).
2. These compounds exhibited correct ir and nmr spectra and satisfactory elemental analysis ($\pm 0.3\%$).
3. For example, see "Ring Index," A.M. Patterson, American Chemical Society, Washington, D.C., 1960, Chapter XI, p. 1046.
4. R.B. Woodward and W.M. McLamore, *J. Amer. Chem. Soc.*, **71**, 379 (1949).